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November 3, 1999

Assistant Commissioner of Patents and Trademarks

Washington, D.C. 20231

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NEW APPLICATION TRANSMITTAL

Transmitted herewith for filing is the patent application of

Inventor: David W. Bergevin

For: Fertilizer Compositions For Administering Phosphates to Plants

Dated: November 1, 1999

The application consists of: 18 pages of specification including 41 claims, a Declaration and Power of Attorney signed by the Inventor, and a check in payment of the filing fee.

Also enclosed is a verified statement that this is a filing by a small entity under 37 CFR 1.9 and 1.27.

The filing fee has been calculated as follows:

Basic fee - Small Entity		\$380.00
Total Claims: 41	=	189.00
Indep. Claims: 3	=	0.00
Dep. Claims: 38	=	0.00
Mult. Dep. Cls.:	=	0.00
FILING FEE	=	\$569.00

A check in the amount of \$569.00 is enclosed to cover the filing fee.

Respectfully submitted,

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cc: David W. Bergevin

Page 1 of 2

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PATENT
File No. P-1274-981

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE

In Re Application of)
DAVID W. BERGEVIN) Our Ref No.: P-1274-991
)
For: FERTILIZER COMPOSITIONS) Date: November 3, 1999
FOR ADMINISTERING)
PHOSPHATES TO PLANTS)

CERTIFICATE OF EXPRESS MAILING

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Dear Sir:

The undersigned hereby certifies that the attached:

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Signature

November 3, 1999

Date

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PATENT
File No. P-1274-981

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE

Applicant or Patentee: David W. Bergevin

Serial or Patent No.: _____

Filed or Issued: _____

For: Fertilizer Compositions For Administering Phosphates to Plants

**VERIFIED STATEMENT (DECLARATION) CLAIMING SMALL ENTITY STATUS
(37 CFR 1.9(f) and 1.27(b) - INDEPENDENT INVENTOR**

As a below named inventor, I hereby declare that I qualify as an independent inventor as defined in 37 CFR 1.9(c) for purposes of paying reduced fees under Section 41(a) and (b) of Title 35, United States Code, to the Patent and Trademark Office with regard to the invention entitled **Fertilizer Compositions For Administering Phosphates to Plants** described in

the specification filed herewith.

application serial no. _____, filed _____

patent no. _____, issued _____

I have not assigned, granted, conveyed or licensed and am under no obligation under contract or law to assign, grant, convey or license, any rights in the invention to any person who could not be classified as an independent inventor under 37 CFR 1.9(c) if that person had made the invention, or to any concern which would not qualify as a small business concern under 37 CFR 1.9(d) or a nonprofit organization under 37 CFR 1.9(e).

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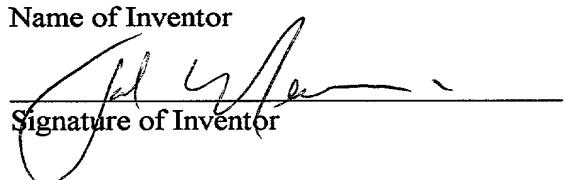
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I acknowledge the duty to file, in this application or patent, notification of any change in status resulting in loss of entitlement to small entity status prior to paying, or at the time of paying, the earliest of the issue fee or any maintenance fee due after the date on which status as a small entity is no longer appropriate. (37 CFR 1.28(b)).

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code, and that such willful false statements may jeopardize the validity of the application, any patent issuing thereon, or any patent to which this verified statement is directed.

David W. Bergevin

Name of Inventor



Date: 11/3/99

Signature of Inventor

306 N. Montana Court, Kennewick, WA 99336

Address of Person Signing

9 Fertilizer Compositions For Administering Phosphates to Plants
10
11

12 Field of the Invention
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15 The present invention relates generally to fertilizer compositions and a method
16 for administering the same. In particular the invention relates to phosphate fertilizer
17 compositions including citric acid and a method of providing phosphates to plants.

18 Background of the Invention
19
20

21 Phosphate is an essential element in plant growth and is common in soil.
22 Phosphates are classified by the percentage of phosphate that is soluble in water and
23 citrate. The sum of the water-soluble and citrate-soluble phosphates (hereinafter
24 referred to as soluble phosphates) is the amount available to plant roots.

25 Soils contain soluble and insoluble phosphates. By far the vast majority of
26 phosphate compounds in soil are insoluble, normally in the form of brushites. Soluble
27 phosphates may react with natural cations in the soil such as Ca^{+1} or Ca^{+2} . Such
28 phosphates become brushites and other insoluble crystalline compounds. Such
29 insoluble phosphates are not available to plant roots. Soluble phosphates can be bio-
30 organic complexes or simple soluble salts called orthophosphates, i.e., salts containing
 PO_4^{-3} , HPO_4^{-2} , or $\text{H}_2\text{PO}_4^{-1}$. The bio-organic complexes are produced from plant and
animal wastes and degradation.

31 Fertilizers utilize orthophosphates and polyphosphates (chains of
32 orthophosphates). However, not all of the orthophosphate is available to the plant
33 roots. When the orthophosphate comes in contact with a strong cation in the soil, an
34 insoluble crystal is formed such as a brushite or monetite. This problem is overcome
35 by saturating the cation rich soil with soluble phosphates. Over saturation of
36 orthophosphates in the soil ties up the inhibiting cations, which then allows some
soluble phosphates to travel from the surface of the soil to the plant roots. However,

1 this practice is wasteful and not environmentally sound. The excess phosphates can
2 run off into waterways where the results can be devastating.

3 Chelating agents have recently been used to chelate and transport metal ions to
4 plant roots. U.S. Pat No. 5,372,626 discloses a method for providing metal ions
5 consisting of Fe^{+3} , Mn^{+2} , Cu^{+2} , and Zn^{+2} to plant roots by using citric acid to chelate
6 the metal ions for transport through the soil. The use of humic acid as a chelating
7 agent to produce a fertilizer containing stabilized predetermined levels of water-
8 soluble and citric-acid soluble phosphoric contents is disclosed in U.S. Pat No.
9 4,588,431.

10 The patents referred to herein are provided herewith in an Information
11 Disclosure Statement in accordance with 37 CFR 1.97.

12 Summary of the Invention

13 The present invention disclosed herein includes a plant fertilizer composition
14 that includes an acid acting as a steric transport vehicle (STV) that may react with,
15 chelate, or block any metal ions in the soil in which the plants grow during
16 administration of the fertilizer. This reaction renders the metal ions substantially
17 ineffective for reacting with the fertilizer components. The fertilizer is thereby
18 shielded from interference with any metal ions in the soil and the delivery of the
19 fertilizer to the plant roots is thereby enhanced.

20 The acid may be, but is not limited to, an organic acid and preferably citric
21 acid.

22 The fertilizer composition may include phosphorus. This means that the
23 fertilizer may include any phosphorus containing compound, acid, or salt thereof.
24 Likewise, in one embodiment, the fertilizer composition may include phosphate ions
25 and citric acid, the citric acid chelating any metal ions in the soil in which the plants
26 grow thereby facilitating the transportation of phosphate to the plant roots.

27 In an alternative embodiment, the fertilizer composition may comprise metal
28 ions, said acid reacting with said any metal ions in said soil in which the plants grow,
29 thereby inhibiting any interference from said any metal ions in said soil in which the
30 plants grow and enhancing delivery of the fertilizer metal ions to the plant roots.

The invention further discloses a method for providing a fertilizer to plants,

1 comprising the steps of administering to the plants a fertilizer composition; and
2 shielding the fertilizer from interference with any compounds including metal ions
3 during said administration of the fertilizer to the plant thereby enhancing delivery of
4 the fertilizer to the plants.

5 In one embodiment, the method comprises the steps of administering to the
6 soil in which the plants grow, a plant soil fertilizer composition; and shielding the
7 fertilizer from interference with any metal ions in the soil in which the plants grow
8 during said administration of the fertilizer to the plant roots, thereby enhancing
9 delivery of the fertilizer to the plant roots.

10 In an alternative embodiment of the method, the fertilizer composition may be
11 foliarly applied, the steric transport vehicle reacting with or loosening up the foliar
12 biochemical constituents to thereby facilitate the uptake of the fertilizer.

13 Detailed Description

14 The present invention disclosed herein includes a plant fertilizer composition
15 that includes an acid, acting as a steric transport vehicle (STV), whereby said acid
16 may react with, chelate, or block any metal ions in the soil in which the plants grow
17 during administration of the fertilizer. This reaction renders the metal ions
18 substantially ineffective for reacting with the fertilizer components. The fertilizer is
19 thereby shielded from interference with any metal ions in the soil and the delivery of
20 the fertilizer to the plant roots is thereby enhanced.

21 The acid may be an organic acid. The organic acid may be, but is not limited
22 to, an aliphatic acid, or an acid containing at least one carboxyl group. The organic
23 acid is preferably citric acid.

24 Likewise, the acid may be, but is not limited to, phosphoric acid, phosphorous
25 acid, an acid with a molecular weight of not more than 400, a phosphorus-containing
26 acid with a molecular weight of not more than 300, a sulfur-containing acid, oxalic
27 acid, and acetic acid. The sulfur-containing acid may be, but is not limited to, sulfuric
28 acid and sulfurous acid.

29 The fertilizer composition may include phosphorus. This means that the
30 fertilizer may include any phosphorus containing compound, acid, or salt thereof. For
example, the fertilizer may include a soluble phosphate compound which may be, but

1 is not limited to, a polyphosphate compound and an orthophosphate compound.
2 Likewise, in one embodiment, the fertilizer composition may include
3 phosphate ions and citric acid, the citric acid chelating any metal ions in the soil in
4 which the plants grow thereby facilitating the transportation of phosphate to the plant
5 roots. It has been experimentally determined that the molar ratio of citric acid
6 concentration to the concentration of the phosphate ions is most preferably about 0.25
7 to 2.0.

8 In an alternative embodiment, the fertilizer composition may comprise metal
9 ions, said acid reacting with said any metal ions in said soil in which the plants grow,
10 thereby inhibiting any interference from said any metal ions in said soil in which the
11 plants grow and enhancing delivery of the fertilizer metal ions to the plant roots.

12 The composition of the present invention may be provided as a dry particulate
13 solid which may be administered to the soil as such or after its dissolution in water.
14 Compositions in accordance with the present invention may also be provided in the
15 form of an aqueous solution either ready for administration or as a concentrate which
16 is diluted prior to use.

17 The invention further discloses a method for providing a fertilizer to plants,
18 comprising the steps of administering to the plants a fertilizer composition; and
19 shielding the fertilizer from interference with any compounds including metal ions
20 during said administration of the fertilizer to the plant thereby enhancing delivery of
21 the fertilizer to the plants.

22 The fertilizer may be administered by various means known per se. The
23 fertilizer may be a dry particulate solid or aqueous solution as described above. In
24 one embodiment, the method may comprise the steps of administering to the soil in
25 which the plants grow, a plant soil fertilizer composition; and shielding the fertilizer
26 from interference with any metal ions in the soil in which the plants grow during said
27 administration of the fertilizer to the plant roots, thereby enhancing delivery of the
28 fertilizer to the plant roots.

29 In an alternative embodiment of the method, the fertilizer composition may be
30 foliarly applied, a steric transport vehicle (STV) reacting with or loosening up the
foliar biochemical constituents to thereby facilitate the uptake of the fertilizer.

1 The fertilizer composition may be shielded by the use of a steric transport
2 vehicle (STV) included in the fertilizer. The STV may use pH, steric blocking,
3 passive cation holdup, and chelation to facilitate the transportation of the fertilizer to
4 the plants.

5 In one embodiment, said fertilizer is shielded from interference with said any
6 metal ions in said soil by chelating said any metal ions.

7 The STV is preferably an acid. Acids are preferred that have a network of
8 carboxyl groups working in unison and synergistically share β -hydroxyl and α -
9 hydroxyl groups. These properties foster the hold up or chelation of cations in the
10 soil. To maximize effectiveness, it is further preferred that the STV be transported to
11 the plants for uptake along with the fertilizer constituents, e.g., the water soluble
12 constituents such as phosphates. Thus, acids, particularly organic acids, of small
13 molecular weights are preferred due to their similar migration and holdup
14 characteristics, as that of the fertilizer constituents including, e.g., phosphate
15 compounds. The organic acid may be, but is not limited to, an aliphatic acid or an
16 acid containing at least one carboxyl group.

17 The organic acid is preferably citric acid. The size to carboxyl site ratio of
18 citric acid, combined with its sharing of α and β carbons makes citric acid particularly
19 effective to chelate metal ions in the soil. Its multiple active sites act in unison to
20 make light common bonds with metals. The β -hydroxyl and α -carboxyl groups on
21 this tricarboxylic acid have been shown to be much more effective to hold up or
22 chelate cations than other acids including simple organic acids, and aromatic acids.

23 The large steric blocking properties of citric acid block out cations and shield the
24 fertilizer constituents, e.g., the phosphates. Citric acid further uses pH to enhance the
25 solubility of the fertilizer constituents, e.g., phosphate compounds. The size of the
26 citrate compound and its vastly negative sites, along with its solubility, allows it to be
27 transported along with the fertilizer constituents, e.g., water soluble constituents such
28 as phosphates, and further allows it to perform an interference against a cation attack.

29 Likewise, the acid may be, but is not limited to, phosphoric acid, phosphorous
30 acid, an acid with a molecular weight of not more than 400, a phosphorus-containing
acid with a molecular weight of not more than 300, a sulfur-containing acid, oxalic

1 acid, and acetic acid. The sulfur-containing acid may be, but is not limited to, sulfuric
2 acid and sulfurous acid.

3 The fertilizer composition may include phosphorus. This means that the
4 fertilizer may include any phosphorus containing compound, acid, or salt thereof. For
5 example, the fertilizer may include a soluble phosphate compound which may be, but
6 is not limited to, a polyphosphate compound and an orthophosphate compound.

7 Likewise, in one embodiment of the method, the fertilizer composition may
8 include phosphate ions and citric acid, the citric acid chelating any metal ions in the
9 soil in which the plants grow thereby facilitating the transportation of phosphate to the
10 plant roots. It has been experimentally determined that the molar ratio of citric acid
11 concentration to the concentration of the phosphate ions is most preferably about 0.25
12 to 2.0.

13 In an alternative embodiment of the method, the fertilizer composition may
14 comprise metal ions, said acid reacting with said any metal ions in said soil in which
15 the plants grow, thereby inhibiting any interference from said any metal ions in said
16 soil in which the plants grow and enhancing delivery of the fertilizer metal ions to the
17 plant roots.

18 In order that the invention described herein may be more fully understood, the
19 following examples are set forth. It should be understood that these examples are for
20 illustrative purposes only and are not to be construed as limiting the scope of the
21 invention in any manner.

22 **EXAMPLE 1.**

23 This example illustrates a method of producing a concentrated STV enhanced
24 phosphate solution (production concentrate) of this invention. The solution was
25 prepared in a 1.2 L reactor equipped with a stirrer. The following reactants were
26 added in order at STP: 700 ml of water, 200 g of monoammonium phosphate, 100 g
27 of 2-hydroxy-1,2,3-propane tricarboxylic acid. Stirring was continued until all solids
28 are dissolved. The resulting STV enhanced phosphate solution contained 20 percent
29 monoammonium phosphate by weight (2.49 M available $H_2PO_4^-$).

30 **EXAMPLE 2.**

1 This example illustrates a method of producing a dilute STV enhanced
2 phosphate solution (application solution) of this invention. The solution was prepared
3 in a 1.2 L reactor equipped with a stirrer. The following reactants were added in
4 order at STP: 700 ml of water, 20 g of monoammonium phosphate, 10 g of 2-
5 hydroxy-1,2,3-propane tricarboxylic acid. Stirring was continued until all solids are
6 dissolved. The resulting STV enhanced phosphate solution contained 2 percent
7 monoammonium phosphate by weight (0.249 M available $H_2PO_4^-$ ions).

EXAMPLE 3.

9 This example illustrates that phosphate STV of this invention mitigates
10 calcium cation interference with phosphate anion availability. An STV enhanced
11 phosphate solution was prepared as described in Example 2. In addition a control was
12 prepared in a manner identical to Example 2, with the omission of the 2-hydroxy-
13 1,2,3-propane tricarboxylic acid. Further, thirteen additional solutions were prepared
14 in a manner identical to Example 2, with the substitution of other acids in place of the
15 2-hydroxy-1,2,3-propane tricarboxylic acid. Resulting in the 15 solutions shown in
16 Table 1. Conc. refers to concentration.

TABLE 1

Solution Sample Number	Phosphate Conc.	Acid	
		Name	Conc.
1 (control)	.25 M H_2PO_4^- ions	None	0 M
2	.25 M H_2PO_4^- ions	2-hydroxy-1,2,3-propane tricarboxylic acid	.074 M
3	.25 M H_2PO_4^- ions	Acetic	.074 M
4	.25 M H_2PO_4^- ions	Boric	.074 M
5	.25 M H_2PO_4^- ions	Fumaric	.074 M
6	.25 M H_2PO_4^- ions	Glycolic	.074 M
7	.25 M H_2PO_4^- ions	Hydrochloric	.074 M
8	.25 M H_2PO_4^- ions	Malic	.074 M
9	.25 M H_2PO_4^- ions	Nitric	.074 M
10	.25 M H_2PO_4^- ions	Nitrous	.074 M
11	.25 M H_2PO_4^- ions	Oxalic	.074 M
12	.25 M H_2PO_4^- ions	Salicylic	.074 M
13	.25 M H_2PO_4^- ions	Sulfuric	.074 M
14	.25 M H_2PO_4^- ions	Sulfurous	.074 M
15	.25 M H_2PO_4^- ions	Tartaric	.074 M

30 A highly enriched calcium cation laden soil was simulated by mixing 1 part

1 gibsome powder, and 1 part 50 mesh white quartz. Three hundred grams of the
2 simulated calcium cation laden soil was slowly added to the 14 samples (shown in
3 Table 1) while they were agitated on a shaker bed. The solutions were allowed to
4 react at STP for 23 hours. At which time the mixtures were allowed to settle for one
5 hour, and then the uppermost 100 ml aliquot of the liquid portions were decanted. A
6 heteropoly blue colorimetric method was used to determine the soluble inorganic
7 phosphate concentration of the decanted aliquots. The results are shown in Table 2.
8 These results are averages of duplicated evolutions expressed as a method specific
9 solubility coefficient described by the equation:

10

$$S_{P_X} = [P_X]/[P_{E2}]$$

11 where,

12 S_{P_X} = The method specific solubility coefficient for soluble inorganic phosphate
13 concentration in the samples shown in Table 1, and treated with calcium as
14 previously described in Example 3.

15 X = The acid used

16 $[P_X]$ = The soluble inorganic phosphate concentration in decanted acid solution
17 aliquot, as determined a heteropoly blue colorimetric method.

18 $[P_{E2}]$ = The soluble inorganic phosphate concentration in the solution prepared in
19 Example 2 (no calcium addition), as determined a heteropoly blue colorimetric
20 method.

21

TABLE 2.

22

Sample	ACID (X)		Properties			
	Number	Name	Descriptive Formula	S_{P_X}	K	pK
2	Citric		$\text{HO}(\text{CH}_2\text{COOH})_2\text{COOH}$	0.615	0.00071	3.14
12	Sulfuric		H_2SO_4	0.536	0.012	1.92
13	Sulfurous		H_2SO_3	0.495	0.0154	1.81

1	11	Oxalic	HOOC-COOH	0.494	.0590	1.23
2	14	Tartaric	HOOCCHOHCHOHCOOH	0.446	0.00104	2.98
3	8	Malic	HOOCCHOHCH ₂ COOH	0.398	0.00039	3.4
4	5	Fumaric	HOOCCH=CHCOOH	0.343	0.00093	3.03
5	3	Acetic	CH ₃ COOH	0.291	1.76E-05	4.75
6	6	Glycolic	HOCH ₂ COOH	0.287	0.000148	3.83
7	9	Nitric	HNO ₃	0.271		
8	11	Salicylic	HOCH ₃ COOH	0.263	3.3E-5	4.48
9	10	Nitrous	HNO ₂	0.25	0.00046	3.37
10	7	Hydrochloric	HCl	0.238		
11	4	Boric	H ₃ BO ₃	0.193	7.3E-10	9.14
12	1	Control (none)	N/A	0.079	N/A	N/A
13						
14						
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20 The resulting method specific solubility coefficient indicate that citric acid is
 21 the most effective STV under these conditions. It also indicates that some inorganic
 22 acids such as sulfuric and sulfurous acids can be very effective. Phosphoric and
 23 phosphorous acids could not be tested with these other acids, because of the
 24 heteropoly blue colorimetric test can be fooled by their soluble phosphor.

25 **EXAMPLE 4.**

26 This example illustrates that the 2-hydroxy-1,2,3-propane tricarboxylic acid
 27 phosphate STV of this invention provide phosphate in a form which can be utilized by
 28 plants. An STV enhanced phosphate solution was prepared as described in Example
 29 2. In addition a control was prepared in a manner identical to Example 2, with the
 30 omission of the 2-hydroxy-1,2,3-propane tricarboxylic acid. Further, two additional

1 solutions were prepared in a manner identical to Example 2, with the substitution of
2 other acids in place of the 2-hydroxy-1,2,3-propane tricarboxylic acid. The four
3 solutions are shown in Table 3.

4 TABLE 3.

5 6 7 8 9 Solution Sample Number	Phosphate Concentration	Acid	
		Name	Concentration
1 (control)	.25 M H_2PO_4^- ions	None	0 M
2	.25 M H_2PO_4^- ions	2-hydroxy-1,2,3-propane tricarboxylic acid	.074 M
3	.25 M H_2PO_4^- ions	phosphoric acid	.074 M
4	.25 M H_2PO_4^- ions	phosphorous acid	.074 M

10 Corn was planted in soil which was deficient in phosphate but adequate in
11 other nutrients. When the corn was approximately 12 inches tall, the four phosphate
12 solutions, shown in Table 3, were foliarly applied. The comparison areas were
13 treated with 0, 5, 10, 15 and 20 gallons of solution per acre.

14 After one week of exposure, samples of the whorl leaves of the corn were
15 taken. These samples were washed and analyzed for plant phosphate level in the
16 tissue. The results, shown in Table 4, indicate that the increase in phosphate uptake at
17 the 10 gallon treatment level was 4677 ppm with 2-hydroxy-1,2,3-propane
18 tricarboxylic acid phosphate STV, and only 696 ppm with the control. As expected
19 the phosphor acids were also shown to be effective phosphate STVs.

20 TABLE 4.

21 22 23 Treatment Level	Phosphate Uptake in Corn (in ppm)			
	24 25 26 27 Gallons of Solution/Acre	Sample 1		Sample 2
		Concentration	Enhancement	Concentration
0	3125	0	3068	0
5	3428	302	4732	1664
10	3821	696	7745	4677
15	4172	1047	8532	5464
20	4376	1251	8957	5889
Sample 3				
0	3107	0	3097	0
5	3920	813	3848	751
10	5346	2239	5625	2528
15	7211	4104	7513	4416

1	20	7538	4431	7569	4472
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EXAMPLE 5.

4 This example illustrates the effect of the molar ratio of 2-hydroxy-1,2,3-
5 propane tricarboxylic acid on the effectiveness of the phosphate STV. This also
6 illustration that the preferred molar ratio of 2-hydroxy-1,2,3-propene tricarboxylic
7 acid to phosphate ions is 1:4 to 2:1. Nine samples were prepared in a manner similar
8 the method described in Example 2. The samples differ only in the 2-hydroxy-1,2,3-
9 propane tricarboxylic acid to phosphate ions molar ratio. The molar ratios of these
10 samples are detailed in Table 5.

11 Corn was planted in soil which was deficient in phosphate but adequate in
12 other nutrients. When the corn was approximately 12 inches tall, the eight phosphate
13 solutions, shown in table 5, were foliarly applied. The comparison areas were each
14 treated with 10 gallons per acre, of one of the solution shown in Table 5.

TABLE 5.

Molar Ratio (Citric Acid:Phosphate ions)	Phosphate Uptake in Corn (in ppm)		
	Moles of Citric Acid/Liter	Moles of Phosphate Ions/Liter	enhancement (ppm) (3068 ppm tare)
(1:8)	0.031 M Citric Acid	.25 M H_2PO_4^- ions	2575
(1:6)	0.042 M Citric Acid	.25 M H_2PO_4^- ions	3073
(1:4)	0.062 M Citric Acid	.25 M H_2PO_4^- ions	4529
(1:2)	0.125 M Citric Acid	.25 M H_2PO_4^- ions	4732
(1:1)	0.250 M Citric Acid	.25 M H_2PO_4^- ions	4821
(2:1)	0.500 M Citric Acid	.25 M H_2PO_4^- ions	4976
(4:1)	1.000 M Citric Acid	.25 M H_2PO_4^- ions	5487
(8:1)	2.000 M Citric Acid	.25 M H_2PO_4^- ions	5264

25 After one week of exposure, samples of the whorl leaves of the corn were
26 taken. These samples were washed and analyzed for plant phosphate level in the
27 tissue. The results, shown in Table 5, indicate that the effective molar ratio is between
28 1:4 and 2:1. Molar ratios below this range are less effective, and higher molar ratios
29 show diminishing returns.

1 While a preferred embodiment of the present invention has been shown and
2 described, it will be apparent to those skilled in the art that many changes and
3 modifications may be made without departing from the invention in its broader
4 aspects. The appended claims are therefore intended to cover all such changes and
5 modifications as fall within the true spirit and scope of the invention.

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1 I CLAIM:

2 1. A method for providing a fertilizer to plant roots, comprising the steps of:

3 a. administering to the soil in which the plants grow, a plant soil fertilizer

4 composition; and

5 b. shielding the fertilizer from interference with any metal ions in the soil

6 in which the plants grow during said administration of the fertilizer to

7 the plant roots, thereby enhancing delivery of the fertilizer to the plant

8 roots.

9 2. The method of claim 1 wherein said fertilizer is shielded from interference with

10 said any metal ions in said soil by chelating said any metal ions.

11 3. The method of claim 1 wherein said fertilizer includes an acid, said acid

12 reacting with said any metal ions in said soil in which the plants grow, thereby

13 inhibiting any interference from said any metal ions in said soil in which the

14 plants grow and enhancing delivery of the fertilizer to the plant roots.

15 4. The method of claim 3 wherein said acid is an organic acid.

16 5. The method of claim 4 wherein said organic acid is selected from the group

17 consisting of an acid containing at least one carboxylic group, an aliphatic

18 acid, and an aliphatic multi-carboxyl acid.

19 6. The method of claim 4 wherein said organic acid is citric acid.

20 7. The method of claim 6 wherein the fertilizer includes phosphate ions and the

21 molar ratio of citric acid concentration to the concentration of the phosphate

22 ions is about 0.125 to 8.0.

23 8. The method of claim 7 wherein said molar ratio is about 0.25 to 4.0.

24 9. The method of claim 7 wherein said molar ratio is about 0.25 to 2.0.

25 10. The method of claim 3 wherein the fertilizer includes phosphorus.

26 11. The method of claim 3 wherein the fertilizer includes a soluble phosphate

27 compound selected from the group consisting of a polyphosphate compound

28 and an orthophosphate compound.

29 12. The method of claim 10 wherein said acid is selected from the group

30 consisting of phosphoric acid, phosphorous acid, an acid with a molecular

1 weight of not more than 400, a phosphorus-containing acid with a molecular
2 weight of not more than 300, sulfuric acid, sulfurous acid, oxalic acid, and
3 acetic acid.

4 13. The method of claim 3 wherein said acid is a sulfur-containing acid.

5 14. The method of claim 3 wherein the fertilizer includes metal ions, said acid
6 reacting with said any metal ions in said soil in which the plants grow, thereby
7 inhibiting any interference from said any metal ions in said soil in which the
8 plants grow and enhancing delivery of the fertilizer metal ions to the plant
9 roots.

10 15. The method of claim 14 wherein said acid is selected from the group
11 consisting of phosphoric acid, phosphorous acid, an acid with a molecular
12 weight of not more than 400, a phosphorus-containing acid with a molecular
13 weight of not more than 300, sulfuric acid, sulfurous acid, oxalic acid, and
14 acetic acid.

15 16. A plant fertilizer composition comprising an acid whereby said acid reacts with
16 any metal ions in the soil in which the plants grow thereby shielding the
17 fertilizer from interference with any metal ions in the soil and enhancing
18 delivery of the fertilizer to the plant roots.

19 17. The composition of claim 16 wherein said acid is an organic acid.

20 18. The composition of claim 17 wherein said organic acid is selected from the
21 group consisting of an acid containing at least one carboxylic group, an
22 aliphatic acid, and an aliphatic multi-carboxyl acid.

23 19. The composition of claim 17 wherein said organic acid is citric acid.

24 20. The composition of claim 19 wherein the fertilizer includes phosphate ions and
25 the molar ratio of citric acid concentration to the concentration of the
26 phosphate ions is about 0.125 to 8.0.

27 21. The composition of claim 19 wherein said molar ratio is about 0.25 to 4.0.

28 22. The composition of claim 19 wherein said molar ratio is about 0.25 to 2.0.

29 23. The composition of claim 16 wherein the fertilizer includes phosphorus.

30

1 24. The composition of claim 16 wherein the fertilizer includes a soluble
2 phosphate compound selected from the group consisting of a polyphosphate
3 compound and an orthophosphate compound.

4 25. The composition of claim 23 wherein said acid is selected from the group
5 consisting of phosphoric acid, phosphorous acid, an acid with a molecular
6 weight of not more than 400, a phosphorus-containing acid with a molecular
7 weight of not more than 300, sulfuric acid, sulfurous acid, oxalic acid, and
8 acetic acid.

9 26. The composition of claim 16 wherein said acid is a sulfur-containing acid.

10 27. The composition of claim 16 wherein the fertilizer comprises metal ions, said
11 acid reacting with said any metal ions in said soil in which the plants grow,
12 thereby inhibiting any interference from said any metal ions in said soil in
13 which the plants grow and enhancing delivery of the fertilizer metal ions to the
14 plant roots.

15 28. The composition of claim 27 wherein said acid is selected from the group
16 consisting of phosphoric acid, phosphorous acid, an acid with a molecular
17 weight of not more than 400, a phosphorus-containing acid with a molecular
18 weight of not more than 300, sulfuric acid, sulfurous acid, oxalic acid, and
19 acetic acid.

20 29. A method for providing a fertilizer to plant foliage, comprising the step of:
21 a. administering to the foliage, a plant soil fertilizer composition that
22 includes an acid, whereby said acid enhances delivery of the fertilizer to
23 the plant.

24 30. The method of claim 29 wherein said acid is an organic acid.

25 31. The method of claim 30 wherein said organic acid is selected from the group
26 consisting of an acid containing at least one carboxylic group, an aliphatic
27 acid, and an aliphatic multi-carboxyl acid.

28 32. The method of claim 30 wherein said organic acid is citric acid.

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1 33. The method of claim 32 wherein the fertilizer includes phosphate ions and the
2 molar ratio of citric acid concentration to the concentration of the phosphate
3 ions is about 0.125 to 8.0.

4 34. The method of claim 33 wherein said molar ratio is about 0.25 to 4.0.

5 35. The method of claim 33 wherein said molar ratio is about 0.25 to 2.0.

6 36. The method of claim 30 wherein the fertilizer includes phosphorus.

7 37. The method of claim 30 wherein the fertilizer includes a soluble phosphate
8 compound selected from the group consisting of a polyphosphate compound
9 and an orthophosphate compound.

10 38. The method of claim 36 wherein said acid is selected from the group
11 consisting of phosphoric acid, phosphorous acid, an acid with a molecular
12 weight of not more than 400, a phosphorus-containing acid with a molecular
13 weight of not more than 300, sulfuric acid, sulfurous acid, oxalic acid, and
14 acetic acid.

15 39. The method of claim 30 wherein said acid is a sulfur-containing acid.

16 40. The method of claim 30 wherein the fertilizer includes metal ions, said acid
17 thereby enhancing delivery of the fertilizer metal ions to the plant roots.

18 41. The method of claim 40 wherein said acid is selected from the group
19 consisting of phosphoric acid, phosphorous acid, an acid with a molecular
20 weight of not more than 400, a phosphorus-containing acid with a molecular
21 weight of not more than 300, sulfuric acid, sulfurous acid, oxalic acid, and
22 acetic acid.

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12 Virginia McIntyre
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6 Abstract of the Disclosure

7 A plant fertilizer composition that includes an acid whereby said acid may react
8 with, chelate, or block any metal ions in the soil in which the plants grow during
9 administration of the fertilizer. This reaction renders the metal ions substantially
10 ineffective for reacting with the fertilizer components. The fertilizer is thereby shielded
11 from interference with any metal ions in the soil and the delivery of the fertilizer to the
12 plant roots is thereby enhanced. The acid may be, but is not limited to, an organic acid
13 and is preferably citric acid. The fertilizer composition may include phosphorus.
14 Likewise, in one embodiment, the fertilizer composition may include phosphate ions and
15 citric acid, the citric acid chelating any metal ions in the soil in which the plants grow
16 thereby facilitating the transportation of phosphate to the plant roots. The invention
17 further discloses a method for providing a fertilizer to plant roots, comprising the steps
18 of administering to the soil in which the plants grow, a plant soil fertilizer composition;
19 and shielding the fertilizer from interference with any metal ions in the soil in which the
plants grow during said administration of the fertilizer to the plant roots, thereby
enhancing delivery of the fertilizer to the plant roots.

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26 Virginia McIntyre
27 (Printed Name)

28 
29 (Signature)

30

**COMBINED DECLARATION AND POWER OF ATTORNEY
FOR PATENT APPLICATION**

As a below named inventor, I hereby declare that:

My residence, post office address, and citizenship are as stated below next to my name.

I believe I am the original, first and sole inventor of the subject matter which is claimed and for which a patent is sought on the invention entitled **Fertilizer Compositions For Administering Phosphates to Plants**, the specification of which

is attached hereto.

was filed on _____ as
Application Serial No. _____

and was amended on _____
(if applicable)

with amendments through _____.
(if applicable)

I hereby state that I have reviewed and understand the contents of the above identified specification, including the claims, as amended by any amendment referred to above.

I acknowledge the duty to disclose information which is material to the examination of this application in accordance with Title 37, Code of Federal Regulations, Sec. 1.56(a).

I hereby claim foreign priority benefits under Title 35, United States Code, Sec. 119 of any foreign application(s) for patent or inventor's certificate listed below and have also identified below any foreign application for patent or inventor's certificate having a filing date before that of the application on which priority is claimed:

no such applications have been filed

such applications have been filed as follows:

Prior Foreign Application(s)
NONE

Priority
Claimed

[] []

(Number) (Country) (Day/Month/Year Filed) Yes No

I hereby claim the benefit under Title 35, United States Code, Sec. 120 of any United States application(s) listed below and, insofar as the subject matter of each of the claims of this application is not disclosed in the prior United States application in the manner provided by the first paragraph of Title 35, United States Code, Sec. 112, I acknowledge the duty to disclose material information as defined in Title 37, Code of Federal Regulations, Sec. 1.56(a) which occurred between the filing date of the prior application and the national or PCT international filing date of this application:

NONE		
(Application Serial No.)	(Filing Date)	Status - patented, pending, abandoned)

I hereby appoint the following attorneys and/or agents to prosecute this application, to file a corresponding international application, and to transact all business in the Patent and Trademark Office connected therewith:

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Kennewick, WA 99336

Direct all phone calls to him/her at (509) 735-3581

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further, that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code and that such willful false statements may jeopardize the validity of the application or any patent issued thereon.

Full name of sole inventor: David W. Bergevin

Inventor's signature



Date

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